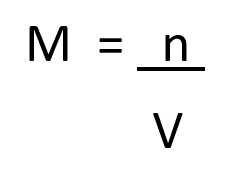
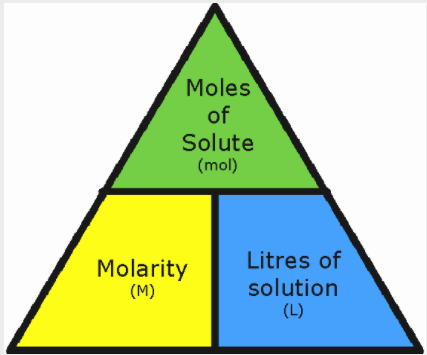
**CW: Molarity Problems**

**Name:**

** **

1. Explain, in words, what molarity means.
2. What is the molarity of a 0.30 L solution containing 0.50 moles of NaCl?
3. Calculate the molarity of 0.289 mol of FeCl3 dissolved in 0.120 L of solution.
4. If a 0.0705 L solution contains 26.00 g of CuCO3, what is the molarity?
5. How many moles of Na2S are in 0.600 L of a 1.55 Molar solution?
6. What mass of H2SO4 is present in 1.63 L of a 0.954 Molar solution?
7. Explain how you would make a 0.60 Molar solution of copper (II) nitrate in a 1.00L container, starting with the mass of copper (II) nitrate you need.