Periodic Table of the Elements

Element Classes

Metal | Nonmetal | Metalloid

Phase at Room Temperature

☐ solid ☐ liquid ☐ gas

alkali metals

alkaline earth metals

Boron Family

Carbon Family

Nitrogen Family

Oxygen Family

Halogens

Noble gases

groups or families are columns

periods → rows (same # orbitals)

	1	2	13	14	15	16	17	18
	1+	2+	3+	4+/-	3-	2-	1-	0
	1e	2e	3e	4e	5e	6e	7e	8e
1	Hydrogen 1 H 1.01							Helium 2 He 4.00
2	Lithium 3 Li 6.94	Beryllium 4 Be 9.01	Boron 5 B 10.81	Carbon 6 C 12.01	Nitrogen 7 N 14.01	Oxygen 8 O 16.00	Fluorine 9 F 19.00	Neon 10 Ne 20.18
3	Sodium 11 Na 22.99	Magnesium 12 Mg 24.31	Aluminum 13 Al 26.98	Silicon 14 Si 28.09	Phosphorous 15 P 30.97	Sulfur 16 S 32.07	Chlorine 17 Cl 35.45	Argon 18 Ar 39.95
4	Potassium 19 K 39.10	Calcium 20 Ca 40.08	Gallium 31 Ga 69.72	Germanium 32 Ge 72.61	Arsenic 33 As 74.92	Selenium 34 Se 78.96	Bromine 35 Br 79.90	Krypton 36 Kr 83.80
5	Rubidium 37 Rb 85.47	Strontium 38 Sr 87.62	Indium 49 In 114.82	Tin 50 Sn 118.71	Antimony 51 Sb 121.76	Tellurium 52 Te 127.60	Iodine 53 I 126.90	Xenon 54 Xe 131.29
6	Cesium 55 Cs 132.91	Barium 56 Ba 137.33	Thallium 81 Tl 204.38	Lead 82 Pb 207.20	Bismuth 83 Bi 208.98	Polonium 84 Po [209]	Astatine 85 At [210]	Radon 86 Rn [222]
7	Francium 87 Fr [223]	Radium 88 Ra [226]	Nihonium 113 Nh [284]	Flerovium 114 Fl [289]	Moscovium 115 Mc [288]	Livermorium 116 Lv [293]	Tennessine 117 Ts [294]	Oganesson 118 Og [294]

Transition Metals (Groups 3 - 12)

	3	4	5	6	7	8	9	10	11	12
4	Scandium 21 Sc 44.96	Titanium 22 Ti 47.88	Vanadium 23 V 50.94	Chromium 24 Cr 52.00	Manganese 25 Mn 54.94	Iron 26 Fe 55.85	Cobalt 27 Co 58.93	Nickel 28 Ni 58.69	Copper 29 Cu 63.55	Zinc 30 Zn 65.39
5	Yttrium 39 Y 88.91	Zirconium 40 Zr 91.22	Niobium 41 Nb 92.91	Molybdenum 42 Mo 95.94	Technetium 43 Tc [98]	Ruthenium 44 Ru 101.07	Rhodium 45 Rh 102.91	Palladium 46 Pd 106.42	Silver 47 Ag 107.87	Cadmium 48 Cd 112.41
6 (57-70)	Lutetium 71 Lu 174.97	Hafnium 72 Hf 178.49	Tantalum 73 Ta 180.95	Tungsten 74 W 183.84	Rhenium 75 Re 186.21	Osmium 76 Os 190.23	Iridium 77 Ir 192.22	Platinum 78 Pt 195.08	Gold 79 Au 196.97	Mercury 80 Hg 200.59
7 (89-102)	Lawrencium 103 Lr [262]	Rutherfordium 104 Rf [261]	Dubnium 105 Db [262]	Seaborgium 106 Sg [266]	Bohrium 107 Bh [264]	Hassium 108 Hs [269]	Meitnerium 109 Mt [268]	Darmstadtium 110 Ds [271]	Roegentium 111 Rg [272]	Copernicium 112 Cn [277]

Inner Transition Metals (Rare Earth Metals)

lanthanides	Lanthanum 57 La 138.91	Cerium 58 Ce 140.12	Praseodymium 59 Pr 140.91	Neodymium 60 Nd 144.24	Promethium 61 Pm [145]	Samarium 62 Sm 150.36	Europium 63 Eu 151.97	Gadolinium 64 Gd 157.25	Terbium 65 Tb 158.93	Dysprosium 66 Dy 162.50	Holmium 67 Ho 164.93	Erbium 68 Er 167.26	Thulium 69 Tm 168.93	Ytterbium 70 Yb 173.04
actinides	Actinium 89 Ac [227]	Thorium 90 Th 232.04	Protactinium 91 Pa 231.04	Uranium 92 U 238.03	Neptunium 93 Np [237]	Plutonium 94 Pu [244]	Americium 95 Am [243]	Curium 96 Cm [247]	Berkelium 97 Bk [247]	Californium 98 Cf [251]	Einsteinium 99 Es [252]	Fermium 100 Fm [257]	Mendeleevium 101 Md [258]	Nobelium 102 No [259]

created by

K Harris

Periodicty (Trends)

1) Valence e^- - electrons in the orbital farthest away from the nucleus. These are the e^-'s involved in chemical reactions (only s & p suborbitals)

Ex) Li ($3e^-$) $1s^2 2s^1$ ← valence suborbital

Na ($11e^-$) $1s^2 2s^2 2p^6 3s^1$ ← valence e^-

K ($19e^-$) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$

☆ Trend as you move down a group is # e^- stays the same ☆

Ex) Li ($3e^-$) $1s^2 2s^1$ 1 valence e^-

Be ($4e^-$) $1s^2 2s^2$ 2 valence e^-

B ($5e^-$) $1s^2 2s^2 2p^1$ 3 valence e^-

C ($6e^-$) $1s^2 2s^2 2p^2$ 4 valence e^-

N ($7e^-$) $1s^2 2s^2 2p^3$ 5 valence e^-

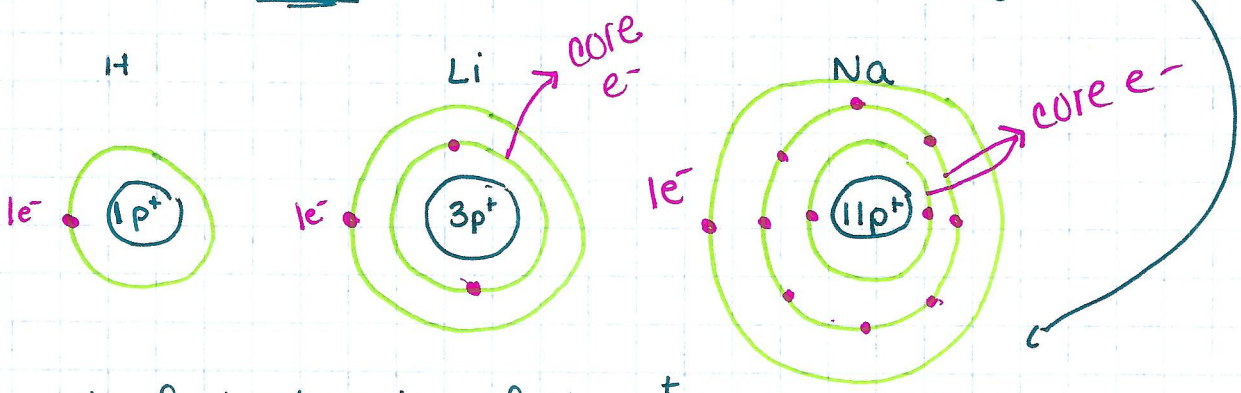
☆ Trend as you move left to right across a period is # valence e^- increases ☆

2) Atomic Radius (size of the atom)

½ the distance b/w the nuclei of 2 identical atoms bonded together

*** Trend - down a group, the radius increases ***

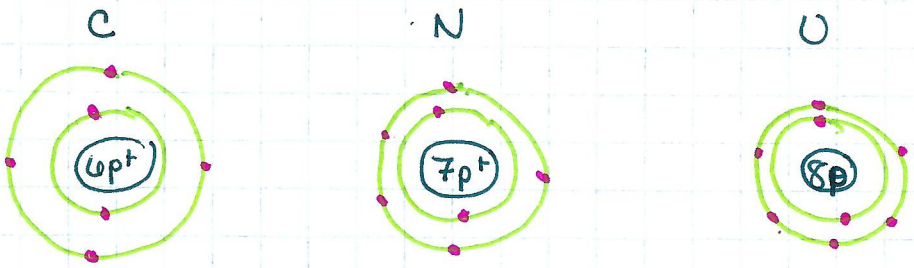
Why? As you go down the group, the elements have more orbitals (core e⁻) b/w the nucleus & the valence e⁻, that shield the valence e⁻ from feeling the



the full attraction of the p^+ in the nucleus so they physically stay farther away making the atoms bigger.

*** Trend - across a period, the radius decreases ***

Why?



4 valence e⁻ 5 valence e⁻ 6 valence e⁻

Shielding remains constant but as you move across a period the #p^+ and #e^- increases which increases the attraction b/w the p^+ & valence e⁻, making them move closer together and the atoms get smaller.

3) 1ST Ionization Energy - energy needed to remove the 1ST valence e^- from an atom.

* down a group - 1ST ionization energy decreases *

Why? The larger the atom is, the more shielding there is b/w the nucleus & valence e^- (less attraction to the nucleus), so less energy is needed to remove the valence e^-

* across a period - 1ST ionization energy increases *

Why? The opposite reason for down a group