

Name: _____

Mixed NOMENCLATURE Practice

Molecular Compounds, Ionic Compounds, & Acids

NAME THE FOLLOWING COMPOUNDS:

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|--|--|
| 1. BaSO ₃ | 16. OF ₂ |
| 2. (NH ₄) ₃ PO ₄ | 17. AgNO ₃ |
| 3. PBr ₅ | 18. As ₂ O ₃ |
| 4. MgSO ₄ | 19. Fe ₂ O ₃ |
| 5. CaO | 20. HClO |
| 6. H ₃ PO ₄ | 21. N ₂ O ₃ |
| 7. Na ₂ Cr ₂ O ₇ | 22. HF |
| 8. MgO | 23. H ₂ C ₂ O ₄ |
| 9. SO ₃ | 24. NaHCO ₃ |
| 10. Cu(NO ₃) ₂ | 25. SiBr ₄ |
| 11. HI | 26. CuCl ₂ |
| 12. N ₂ O | 27. HNO ₂ |
| 13. MnO | 28. SnO ₂ |
| 14. C ₃ H ₈ | 29. BaCrO ₄ |
| 15. H ₃ P | 30. SnO |

WRITE FORMULAS FOR THE FOLLOWING COMPOUNDS:

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| 31. hydrobromic acid | 46. diphosphorus pentoxide |
| 32. chromium(III) carbonate | 47. sulfurous acid |
| 33. magnesium sulfide | 48. lead(II) nitrate |
| 34. iodine trichloride | 49. dihydrogen monoxide |
| 35. lithium hydride | 50. sodium oxalate |
| 36. ammonium hydroxide | 51. perchloric acid |
| 37. calcium chloride | 52. chlorous acid |
| 38. hydroselenic acid | 53. silicon dioxide |
| 39. iron(II) nitride | 54. carbonic acid |
| 40. aluminum hydroxide | 55. sodium chlorate |
| 41. tin(II) fluoride | 56. carbon tetraiodide |
| 42. sulfur tetrachloride | 57. palladium(III) sulfite |
| 43. mercury(II) iodide | 58. xenon hexafluoride |
| 44. dicarbon dichloride | 59. nickel ^(II) nitrate |
| 45. nitric acid | 60. potassium perchlorate |