**CW: Acids Bases: Arrhenius & Brönsted-Lowry Theories Worksheet**

**Name:**

1. How do the Arrhenius and Bronsted-Lowry definitions of an acid and a base differ? How are they similar?
2. What is an amphoteric species? Name one and write a balanced equation that shows why it’s amphoteric.
3. Give the formula for the conjugate base for these acids:
	1. HCl
	2. H2CO3:
	3. H2O:
4. Give the formula of the conjugate acid for these bases:
	1. O2-2
	2. SO42-
	3. H2O:
5. In each equation, label the acids, bases and conjugate pairs
	1. HClO4(aq) + H2SO4(aq) ⇋ ClO4-  + H2SO4+

* 1. NH3 + H3PO4 ⇋ NH4+ + H2PO4-
	2. H2O + HC2H3O2  ⇋ H3O+ + C2H3O2-
1. List the 7 strong acids. Why are they considered strong? What makes an acid weak? What are the strong bases?
2. List the properties of an acid. List the properties of a base.