Atomic Mass Practice

**A.M. = (% isotope 1 × mass isotope 1) + (% isotope 2 × mass isotope 2) +**

**(% isotope 3 × mass isotope 3) + …**

1. Naturally occurring europium (Eu) consists of two isotopes was a mass of 151 and 153. Europium-151 has an abundance of 48.03% and Europium-153 has an abundance of 51.97%. What is the atomic mass of europium?
2. Strontium consists of four isotopes with masses of 84.0 (abundance 0.500%), 86.o (abundance of 9.90%), 87.0 (abundance of 7.00%), and 88.0 (abundance of 82.60%). Calculate the atomic mass of strontium.
3. Titanium has five common isotopes: 46.0Ti (8.00%), 47.0Ti (7.80%), 48.0Ti (73.40%), 49.0Ti (5.50%), 5.00Ti (5.30%). What is the average atomic mass of titanium?

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