Bonding and Nomenclature Study Guide

Tutoring is available after school Thursday and before school by appointment. You can also go to this website for assistance as well.

<http://misterguch.brinkster.net/ionic.html>

<http://misterguch.brinkster.net/covalentcompounds.html>

<http://misterguch.brinkster.net/lewisstructures.html>

**Type III Covalent Compounds**

1. oxygen difluoride
2. sulfur hexafluoride
3. heptane
4. SiF4
5. C5H12
6. NO2

**Type I Regular Ionic Compounds**

1. sodium fluoride
2. potassium sulfide
3. barium cyanide
4. magnesium nitrate
5. ammonium phosphide
6. KCl
7. Na2O
8. RaClO2
9. Na2CO3
10. Cs2SO4

**Type II Transition Metal Ionic Compounds**

1. copper (I) oxide
2. copper (II) oxide
3. iron (III) sulfite
4. lead (IV) hydroxide
5. tungsten (VI) phosphate
6. Cu2S
7. FeO
8. MoF2
9. Fe2(C2O4)3
10. PbS

**Acids**

1. phosphoric acid
2. carbonic acid
3. hydrosulfuric acid
4. hydroiodic acid
5. chlorous acid
6. H2C2O4
7. HClO
8. H2SO4
9. HBr
10. H3P

**Hodgepodge: These can by any type.**

1. carbon dioxide
2. potassium cyanide
3. nitrous acid
4. copper (II) phosphate
5. disilicon trioxide
6. hydrochloric acid
7. aluminum iodide
8. BrF3
9. Li2CO3
10. Fe3(PO3)2
11. HNO3
12. CaF2
13. HF
14. Cu2CO3

**Ionic Bonding: Lewis Structures (Show transfer of electrons and charges formed.)**

1. Li and S
2. Mg and P

**Covalent Bonding: Lewis Structures (Show the sharing of electrons.)**

1. H2O
2. SiF4
3. SO3
4. Review from previous tests.
	1. Write electron configurations for:
		1. Fe
		2. Si
		3. Rb
	2. How many protons, electrons, and neutrons are in
		1. 9140Zr
		2. 22789Ac